Exercise 7.4

EM Radiation & Bohr's Model

Name: _____

Date: _____ Per: _____

The Planck constant (denoted *h*), also called Planck's constant, is a physical constant used to describe the sizes of quanta in quantum mechanics. It is named after Max Planck, one of the founders of quantum theory. The Planck constant is the proportionality constant between energy(*E*) of a photon and the frequency of its associated electromagnetic wave (*v*). This relation between the energy and frequency is called the Planck relation or the Planck–Einstein equation:

$$E = hv$$

Using the following simple relation between frequency (*v*), speed of light (*c*), and wavelength (λ),

$$v = \frac{c}{\lambda}$$

the Planck relation becomes the following

$$E = \frac{hc}{\lambda}$$

$c = 3.00 \times 10^{8}$ m/s
$h = 6.626\ 2 \times 10^{-34}$ J•s (or J/Hz)
1 eV = 1.602×10^{-19} J

DIRECTIONS: Answer the following in the space provided.

1. What is the wavelength of a wave having a frequency of 3.76×10^{14} s^{-1}?
2. What is the frequency of a 6.9×10^{-13} m wave?
3. What is the wavelength of a 2.990 MHz wave?
4. What is the wavelength of a 1.28×10^{17} Hz wave?
5. What is the frequency of a 2600cm wave?
6. What is the wavelength of 109.60 MHz wave?
7. What is the energy of a 7.66×10^{14} Hz wave?
8. What is the frequency of a wave carrying 8.35×10^{-18} J of energy?
9. What is the energy of a 3.12×10^{18} s^{-1} wave?
10. What is the frequency of a 1.310×10^{-22} J wave?
11. What is the energy of a 9330cm wave?
12. What is the wavelength of a 1.32×10^{-6} eV wave?
13. What is the wavelength of a 1.528×10^{-13} J wave?
14. What is the energy in electron-volts (eV) of a 4.22 μm wave?

Revised: 2019-01-29

Answers: 1) 7.98×10^{-7} m 2) 4.3×10^{20} Hz 3) 100.3 m 4) 2.34×10^{-9} m 5) 1.2×10^{7} Hz 6) 2.7372 m 7) 5.08×10^{-19} J 8) 1.26×10^{16} Hz 9) 2.07×10^{-15}J 10) 1.977×10^{11} Hz 11) 2.13×10^{-27} J 12) 0.940 m 13) 1.301×10^{-12} m 14) 0.294 eV

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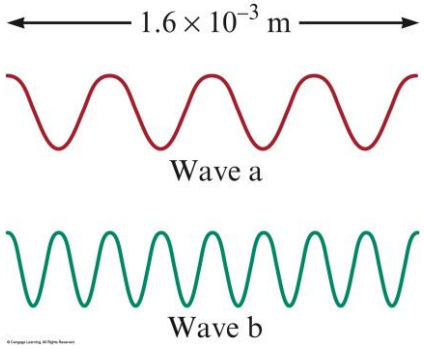
15. Consider the following waves representing electromagnetic radiation:



a. Which wave has the longer wavelength? Calculate the wavelength. _____

b. Which wave has the higher frequency and larger photon energy? Calculate these values.

c. Which wave has the greater velocity? _____

16. According to the Bohr model of the hydrogen atom, how are atomic spectra produced? _____

_____

_____

_____

17. It takes 7.21 x 10^{-19} J of energy to remove an electron from an iron atom. What is the maximum wavelength of light that can do this?

5. Which of the following transitions would produce a photon with the greatest energy? Why?

a. $n = 1 \rightarrow n = 5$

b. $n = 4 \rightarrow n = 3$

c. $n = 5 \rightarrow n = 2$

d. $n = 3 \rightarrow n = 4$

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