



# Exercise 14.2

## Acid/Base Properties

Name: \_\_\_\_\_

Date: \_\_\_\_\_ Per: \_\_\_\_\_

5. How would you prepare 2000 mL of a pH = 1.50 solution using concentrated (12 M) HCl? (A 2 L volumetric flask is available.)

\_\_\_\_\_

\_\_\_\_\_

\_\_\_\_\_

\_\_\_\_\_

6. How would you prepare 2000 mL of a pH = 8.75 solution using solid NaOH? (A 2 L volumetric flask is available.)

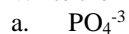
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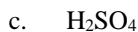
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\_\_\_\_\_

7. Write the formula of the conjugate acid for each of the following bases:



8. Write the formula of the conjugate base for each of the following acids:



9. For each of the following reactions identify the Bronsted-Lowry acid, base, conjugate acid, and conjugate base.

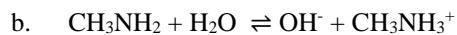


acid: \_\_\_\_\_

base: \_\_\_\_\_

conjugate base: \_\_\_\_\_

conjugate acid: \_\_\_\_\_



acid: \_\_\_\_\_

base: \_\_\_\_\_

conjugate base: \_\_\_\_\_

conjugate acid: \_\_\_\_\_