

## Exercise 6.3b

### Ionic Names to Formulas – Answers

Name: \_\_\_\_\_

Date: \_\_\_\_\_ Per: \_\_\_\_\_

**DIRECTIONS:** Write the correct empirical formula for each of the following.

1. Sodium nitrate	<u>NaNO<sub>3</sub></u>	28. Manganese (VII) oxide	<u>Mn<sub>2</sub>O<sub>7</sub></u>
2. Aluminum sulfate	<u>Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub></u>	29. Mercury (I) oxide	<u>Hg<sub>2</sub>O</u>
3. Potassium chromate	<u>K<sub>2</sub>CrO<sub>4</sub></u>	30. Iron (II) chloride	<u>FeCl<sub>2</sub></u>
4. Lithium hydroxide	<u>LiOH</u>	31. Copper (II) chloride	<u>CuCl<sub>2</sub></u>
5. Ammonium phosphate	<u>(NH<sub>4</sub>)<sub>3</sub>PO<sub>4</sub></u>	32. Tin (II) chloride	<u>SnCl<sub>2</sub></u>
6. Magnesium nitrate	<u>Mg(NO<sub>3</sub>)<sub>2</sub></u>	33. Titanium (IV) iodide	<u>TiI<sub>4</sub></u>
7. Sodium monohydrogen phosphate	<u>Na<sub>2</sub>HPO<sub>4</sub></u>	34. Nickel (II) fluoride	<u>NiF<sub>2</sub></u>
8. Sodium bicarbonate	<u>NaHCO<sub>3</sub></u>	35. Manganese (IV) oxide	<u>MnO<sub>2</sub></u>
9. Potassium bisulfite	<u>KHSO<sub>3</sub></u>	36. Lead (IV) oxide	<u>PbO<sub>2</sub></u>
10. Hydrogen acetate	<u>HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub></u>	37. Nickel (II) hydroxide	<u>Ni(OH)<sub>2</sub></u>
11. Lead (II) chromate	<u>PbCrO<sub>4</sub></u>	38. Mercury (II) oxide	<u>HgO</u>
12. Copper (II) carbonate	<u>CuCO<sub>3</sub></u>	39. Cobalt (III) oxide	<u>Co<sub>2</sub>O<sub>3</sub></u>
13. Silver nitrate	<u>AgNO<sub>3</sub></u>	40. Copper (II) chlorate	<u>Cu(ClO<sub>3</sub>)<sub>2</sub></u>
14. Ferrous sulfide	<u>FeS</u>	41. Cobalt (III) sulfate	<u>Co<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub></u>
15. Manganous bromide	<u>MnBr<sub>2</sub></u>	42. Manganese (III) hydroxide	<u>Mn(OH)<sub>3</sub></u>
16. Barium sulfite	<u>BaSO<sub>3</sub></u>	43. Iron (III) nitrate	<u>Fe(NO<sub>3</sub>)<sub>3</sub></u>
17. Sodium iodide	<u>NaI</u>	44. Tin (IV) nitrate	<u>Sn(NO<sub>3</sub>)<sub>4</sub></u>
18. Calcium nitrate	<u>Ca(NO<sub>3</sub>)<sub>2</sub></u>	45. Cobalt (II) perchlorate	<u>Co(ClO<sub>4</sub>)<sub>2</sub></u>
19. Zinc hydroxide	<u>Zn(OH)<sub>2</sub></u>	46. Chromium (III) sulfate	<u>Cr<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub></u>
20. Potassium oxalate	<u>K<sub>2</sub>C<sub>2</sub>O<sub>4</sub></u>	47. Iron (II) hydroxide	<u>Fe(OH)<sub>2</sub></u>
21. Cuprous chloride	<u>CuCl</u>	48. Copper (II) phosphate	<u>Cu<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub></u>
22. Manganese (III) chloride	<u>MnCl<sub>3</sub></u>	49. Mercury (I) nitrite	<u>Hg<sub>2</sub>(NO<sub>2</sub>)<sub>2</sub></u>
23. Iron (III) bromide	<u>FeBr<sub>3</sub></u>	50. Lead (II) nitrate	<u>Pb(NO<sub>3</sub>)<sub>2</sub></u>
24. Chromium (III) bromide	<u>CrBr<sub>3</sub></u>	51. Mercury (I) sulfate	<u>Hg<sub>2</sub>SO<sub>4</sub></u>
25. Tin (IV) chloride	<u>SnCl<sub>4</sub></u>	52. Tin (IV) oxide	<u>SnO<sub>2</sub></u>
26. Chromium (III) oxide	<u>Cr<sub>2</sub>O<sub>3</sub></u>	53. Manganese (II) bromide	<u>MnBr<sub>2</sub></u>
27. Lead (II) hydroxide	<u>Pb(OH)<sub>2</sub></u>	54. Lithium oxide	<u>Li<sub>2</sub>O</u>