

Exercise 6.5d

Lewis Structures & Exceptions to the Octet Rule

Name: _____

Date: _____ Per: _____

DIRECTIONS: For each of the formulas listed below, draw the Lewis structure and fill in the name of the compound.

Ionic Compounds			
1. AlCl_3	2. Na_2O	3. PbCl_4	4. SnO_2
Name:	Name:	Name:	Name:

Simple Molecules			
5. CS_2	6. CF_4	7. PI_3	8. HCN
Name:	Name:	Name:	Name:

Partial Octets			
9. BeH_2	10. BCl_3	11. BH_3	12. BeCl_2
Name:	Name:	Name:	Name:

Define a partial octet. _____

Which three elements are capable of forming partial octets? _____

Expanded Octets			
13. KrF_2	14. PCl_5	15. SF_6	16. XeF_4
Name:	Name:	Name:	Name:

Define an expanded octet. _____

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At what point in the periodic table do elements gain the ability to form expanded octets? _____

Free Radicals

17. NO	18. NO ₂	19. CH ₃	20. ClO ₂
Name:	Name:	Name:	Name:

Define and describe a free radical. _____

Mixed Review

21. BrO ₂	22. PF ₅	23. CaCl ₂	24. BeI ₂
Name:	Name:	Name:	Name:
25. BI ₃	26. CN ⁻	27. NH ₂ Cl	28. KrCl ₂
Name:	Name:	Name:	Name:
29. NH ₃	30. GeF ₄	31. Li ₂ O	32. PCl ₃
Name:	Name:	Name:	Name: