

Exercise 6.5b

Naming Molecular Compounds

Name: _____

Date: _____ Per: _____

DIRECTIONS: Next to each formula write the name of the covalent compound. For acids write both the molecular name and acid name.

Formulas in *italics* are acids when dissolved in water. The acid name would be used when “(aq)” (meaning aqueous solution) follows the formula. All other times, the molecular name would be used.

Molecular Naming Prefixes		
1 – mono-	5 – penta-	9 – nona-
2 – di-	6 – hexa-	10 – deca-
3 – tri-	7 – hepta-	11 – undeca-
4 – tetra-	8 – octa-	12 – dodeca-

Naming Acids	
Anion suffix	Acid Name
-ide	hydro_____ic acid
-ate	_____ic acid
-ite	_____ous acid

1. OF_2 : _____

2. N_2O_4 : _____

3. PCl_3 : _____

4. CBr_4 : _____

5. *HCl*: _____

Acid: _____

6. *H₂S*: _____

Acid: _____

7. SCl_6 : _____

8. N_2O_5 : _____

9. *HF*: _____

Acid: _____

10. *HClO₃*: _____

Acid: _____

11. NH_3 : _____

12. *HNO₃*: _____

Acid: _____

13. *HNO₂*: _____

Acid: _____

14. H_2O : _____

15. NI_3 : _____

16. *H₂Te*: _____

Acid: _____

17. SiF_4 : _____

18. *H₂Se*: _____

Acid: _____

19. P_2Cl_5 : _____

20. SiO_2 : _____

21. CO_2 : _____

22. *HClO₄*: _____

Acid: _____

23. N_2 : _____

24. *H₂SO₄*: _____

Acid: _____

25. PH_3 : _____

26. *HCN*: _____

Acid: _____

27. *HClO*: _____

Acid: _____

28. C_2H_4 : _____

29. C_2H_2 : _____

30. P_2O_5 : _____